

# Evaluation of the Magnesium Wastes with Boron Oxide in Magnesium Borate Synthesis

A. S. Kipcak, F. T. Senberber, E. Moroydor Derun, S. Piskin

**Abstract**—Magnesium wastes and scraps, one of the metal wastes, are produced by many industrial activities, all over the world. Their growing size is becoming a future problem for the world. In this study, the use of magnesium wastes as a raw material in the production of the magnesium borate hydrates are aimed. The method used in the experiments is hydrothermal synthesis. The conditions are set to, waste magnesium to  $B_2O_3$ , 1:3 as a molar ratio. Four different reaction times are studied which are 30, 60, 120 and 240 minutes. For the identification analyses X-Ray Diffraction (XRD), Fourier Transform Infrared Spectroscopy (FT-IR) and Raman spectroscopy techniques are used. As a result at all the reaction times magnesium borate hydrates are synthesized and the most crystalline forms are obtained at a reaction time of 120 minutes. The overall yields of the production are found between the values of 65-80 %.

**Keywords**—Hydrothermal synthesis, magnesium borates, magnesium wastes, boron oxide

## I. INTRODUCTION

THE increase in the rate of waste formation will be a problem for future of the world due to storage and disposal of wastes. In recent years, studies about waste storage and evaluation are done to solve this problem. Some kinds of waste are hazardous for people and environment. One of the dangerous wastes is metal wastes and scraps. Distribution of waste per person in the United States and Turkey are 8.9% and 7% of metal waste, respectively [1].

Evaluation of wastes is getting more interesting for industrial activities more than storage because of the increasing cost of raw materials and energy. One of the evaluation method is recycle process. Steps of metal waste recycling process are a separate collection, magnetic extraction, smelting factory, recycled metal, comprises the steps of the new metal products. Metal wastes are 1% of solid wastes and 3% of the recovered waste [2]. Magnesium wastes and scraps are produced by many industrial activities, all over the world. Magnesium is a chemical element with the symbol Mg, atomic number 12, and common oxidation number +2.

Elemental magnesium is a fairly strong, silvery-white, light-weight metal (two thirds the density of aluminum). It tarnishes slightly when exposed to air.

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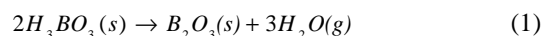
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It is an alkaline earth metal and the eighth most abundant element in the Earth's crust and ninth in the known universe as a whole Magnesium is the fourth most common element in the Earth as a whole (behind iron, oxygen and silicon), making up 13% of the planet's mass and a large fraction of planet's mantle [3]-[5].

Boron is the chemical element with atomic number 5 and the chemical symbol B. The atomic mass is 10.81. It is a low-abundance element in both the solar system and the Earth's crust. However, boron is concentrated on Earth by the water-solubility of its more common naturally occurring compounds, the borate minerals such as boric acid and borax. Turkey has 72.2% of the boron minerals that are present in the world, which can sufficiently meet world's needs for many years. Approximately world's total reserves are reported to be 1.2 billion tons [6], [7].

Boric acid, also called hydrogen borate having chemical formula  $H_3BO_3$ , is a weak acid of boron often used as an antiseptic, insecticide, flame retardant, as a neutron absorber, and as a precursor of other chemical compounds. It exists in the form of colorless crystals or a white powder and dissolves in water. Sometimes written  $B(OH)_3$ . When occurring as a mineral, it is called sassolite [8]. Boron oxide can easily be formed from the following reaction:



Magnesium borate production is useful because of minerals superior properties of thermal and mechanic characters. Main properties of magnesium borate minerals are high heat resistance, corrosion resistance and high coefficient of elasticity. So this boron group can be used in ceramic industry, in the production of superconducting materials, in the composition of detergents, due to the content of boron in the friction-reducing additives in oils and insulating coating compositions [9].

Magnesium borates can be synthesized by liquid-state or solid-state methods. In literature, synthesized magnesium borate minerals with liquid-state method can be listed  $MgBO_2(OH)$  [10],  $MgO \cdot 3B_2O_3 \cdot 17H_2O$  [11],  $MgO \cdot 3B_2O_3 \cdot 3,5H_2O$  [12],  $2MgO \cdot 2B_2O_3 \cdot MgCl_2 \cdot 14H_2O$  [13],  $2MgO \cdot B_2O_3 \cdot H_2O$  and  $MgO \cdot 3B_2O_3 \cdot 7H_2O$  [14]. Synthesized magnesium borate minerals with solid-state method can be listed  $Mg_2B_2O_5$  [15]-[17],  $Mg_3B_2O_6$  [18], [19]. The common feature of all studies done as a raw material MgO or MgOH is to use in synthesis.

In this study, magnesium wastes are used as magnesium source in magnesium borate production. It can be called as a new perspective in evaluation of metal wastes.

## II. EXPERIMENTAL

### A. Raw Material Preparation and Characterization

Magnesium wastes (Fig. 1 (a)) are retrieved from local gold factory in Turkey where these wastes are formed at the instance of plastic molding in the manufacturing processes. These wastes are to be stored in the factory.

Boron oxide is retrieved from the Boron Management Plant in Eskisehir, Turkey. These boron oxides are grinded with agate mortar (Fig. 1 (b)) and sieved (Fig. 1 (c)) to a particle size below 75 microns.

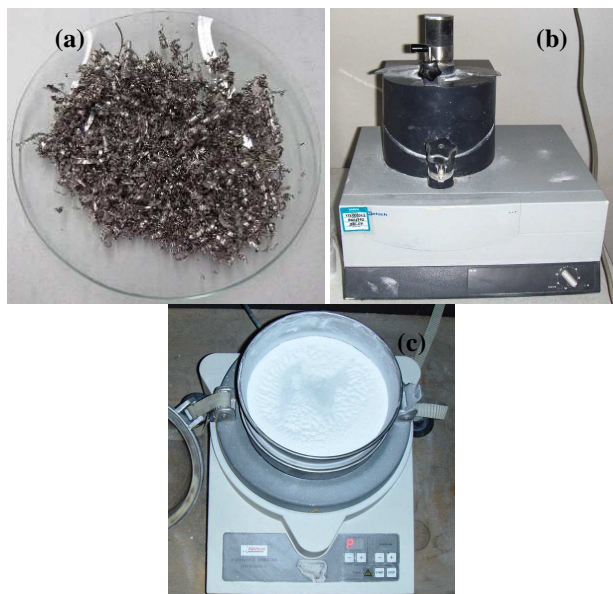


Fig. 1 (a) Waste Mg, (b) Grinding process, (c) Sieving process

Magnesium wastes and boron oxide are subjected to X-Ray Diffraction (XRD) analysis with Philips PANanalytical brand (Fig. 2 (a)) where in this equipment X-rays are produced from Cu-K $\alpha$  tube at the parameters of 45kV and 40mA [6].

Magnesium wastes are subjected to X-Ray Fluorescence (Fig. 2 (b)) analysis by Philips PANanalytical brand Minipal Model 4 with silicon drift detector [20].



Fig. 2 (a) Philips PANanalytical XRD, (b) Philips PANanalytical XRF

Also magnesium wastes are analyzed with Scanning electron microscope with Energy Disperse Spectroscopy (SEM-EDS) by CamScan Apollo 300 field-emission SEM (Fig. 3) and EDS detector brand is Oxford.

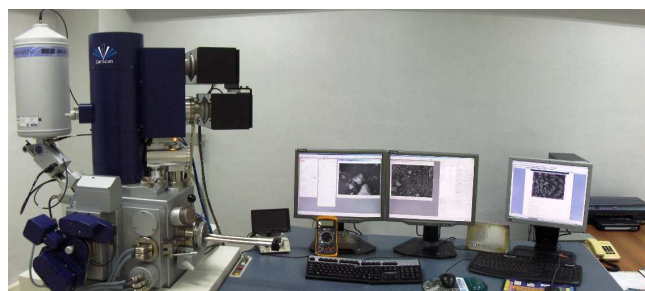


Fig. 3 CamScan Apollo 300 SEM

Boron oxide is subjected to Perkin Elmer Spectrum One (Fig. 4 (a)) Fourier Transform Infrared Spectroscopy (FT-IR) and Perkin Elmer Brand, Raman Station 400F Raman Spectroscopy (Fig. 4 (b)) techniques. In the FT-IR technique Universal ATR sampling accessory – Diamond / ZnSe is used and measurement range is selected as 4000–650 cm<sup>-1</sup>, scan number is 4 and resolution set as 4 cm<sup>-1</sup>. For Raman Spectroscopy, the parameters of exposure time (seconds) and number of exposures was set to 4. Measurement range is selected as 3280–250 cm<sup>-1</sup> and data interval was selected as 2 cm<sup>-1</sup> [20].

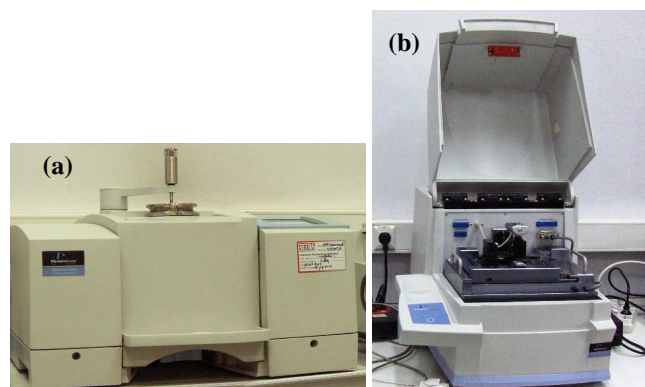
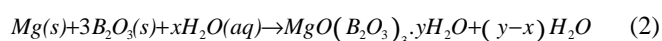


Fig. 4 (a) Perkin Elmer Spectrum One FT-IR, (b) Raman Station 400F Raman Spectroscopy

### B. Hydrothermal Synthesis of Magnesium Borates

For the hydrothermal synthesis, starting molar ratio of waste magnesium to boron oxide is selected as 1:3. The liquid phase is used as ultra-pure water (18.3 m $\Omega$ .cm) that is produced from the equipment of Human Power I+ Water Purification System.

Experiment temperature is selected as 80°C, and four different reaction times are conducted to investigate the phase transition between different types of magnesium borates according to the reaction time changes. The reaction times are set to 30, 60, 120 and 240 minutes. The formation reaction of magnesium borates are shown in (2).



After the reaction, the first filtration process is used for the removal of excess magnesium and other trace amount of metals inside the waste magnesium. In this process, 80°C hot water is used for the washing and dispersing the synthesized magnesium borates below the filter paper. After that the slurry content is dried in Ecocell model oven at 40°C. The dried content was washed and filtered with pure alcohol (96°), supplied from Merck Chemicals, in order to remove excess boric acid content that is formed in the hydrothermal reaction as a byproduct. The byproduct reaction is shown in (3). Then the filtered content was dried in Ecocell model oven again at 40°C.



### C. Characterization of the Synthesized Magnesium Borates

At this step synthesized materials are subjected to XRD, FT-IR and Raman techniques with the parameter set explained at part II.A. "Raw Material Preparation and Characterization".

## III. RESULTS AND DISCUSSION

### A. Raw Material Characterization Results

XRD patterns of the waste magnesium and boron oxide are shown in Fig. 5 and Fig.6 respectively.

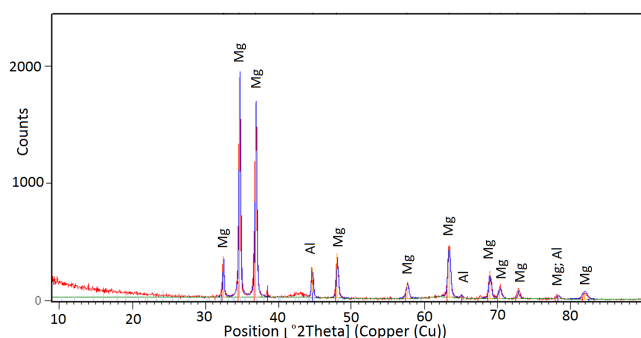


Fig. 5 XRD pattern of waste magnesium [21]

From the waste magnesium XRD pattern, the major peaks represents the 01-089-5003 numbered powder diffraction file (pdf) magnesium also some aluminum minor peaks are observed with pdf number of 01-089-2769 [21].

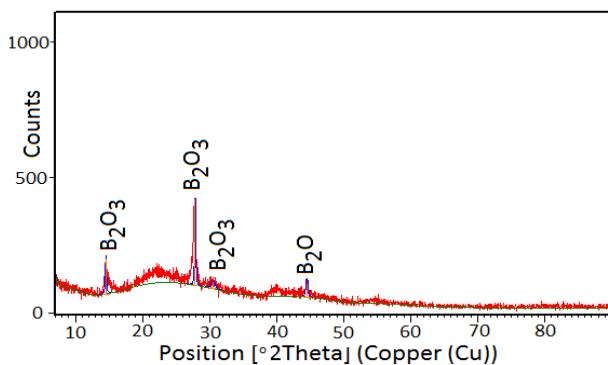


Fig. 6 XRD pattern of boron oxide

From the boron oxide XRD pattern, two different types of "boron oxide" minerals are found. Major peaks represent the "B<sub>2</sub>O<sub>3</sub>" formulated and pdf numbered "00-006-0297" boron oxide and minor peak represents the "B<sub>2</sub>O" formulated and pdf numbered "01-088-2485" boron oxide.

XRF and SEM-EDS results of the waste magnesium is shown in Table I.

TABLE I  
 XRF AND SEM-EDS RESULTS OF THE WASTE MAGNESIUM [21]

Elements	XRF Content (%)	SEM-EDS Content (%)
Mg	93.30	93.12
Al	3.67	3.54
Zn	0.88	1.72
Mn	0.90	1.02
S	0.08	0.21
Ca	0.11	0.14
Cr	0.03	-
Fe	0.93	-
Cu	0.14	0.25

Both XRF and SEM-EDS results showed that the element in the waste magnesium is "magnesium" and the minor element is "aluminum". Other elements can be classified as trace elements. XRF and SEM-EDS analysis supports the XRD analysis.

FT-IR and Raman spectrums of the boron oxide are shown in Fig. 7 and Fig. 8 respectively.

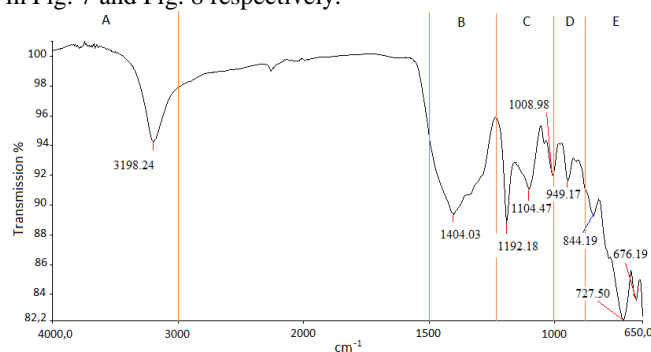


Fig. 7 FT-IR spectrum of the boron oxide

According to the FT-IR analysis of the boron oxide, peak intervals can be classified in five groups. Part A can be explained with the OH groups of the mineral that can be attached to mineral from the humidity of the air. Part B shows the asymmetric stretching of tri-coordinate boron (B<sub>(3)</sub>-O). In part C, the peak peaks between 1192.18 and 1008.98 mean asymmetric stretching of four-coordinate boron (B<sub>(4)</sub>-O). Symmetric stretching of B<sub>(3)</sub>-O can be seen in Part D and at the end of the spectrum, Part E explains the out-of-plane OH<sup>-1</sup> bending band and bending of B<sub>(3)</sub>-O in the structure.

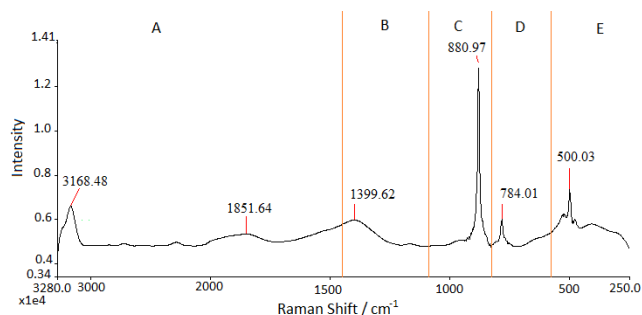


Fig. 8 Raman spectrum of the boron oxide

According to Raman spectrums of  $B_2O_3$ , the first and second peaks in part A might be the  $OH^-$  groups in the minerals that can be attached to mineral from the humidity of the air. The peaks of asymmetric stretching of  $B_{(4)}-O$  is seen in part B. In part C, the peak can be explain with the symmetric stretching of  $B_{(3)}-O$ . In part D symmetric stretching of  $B_{(4)}-O$  is seen. Peak on part E, might be explain with the bending of  $B_{(3)}-O$ .

### B. Synthesized Magnesium Borate XRD Results

XRD patterns and results of the synthesized magnesium borates are shown in Fig. 9 and Table II respectively.

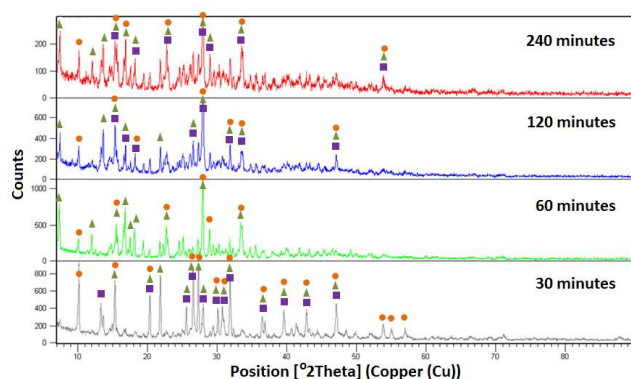


Fig. 9 XRD patterns of the synthesized magnesium borates;  
● Mcallisterite, ▲ Admontite, ■ Magnesium borate hydrate

TABLE II  
XRD RESULTS OF THE SYNTHESIZED MAGNESIUM BORATES

Reac. Time	Pdf #	Mineral Name	Mineral Formula	Score
30	01-076-0540	Admontite	$MgO(B_2O_3)_3 \cdot 7H_2O$	26
	01-070-1902	Mcallisterite	$Mg_2((B_6O_7)(OH)_6)_2 \cdot 9H_2O$	82
	01-073-0638	MgBH*	$MgB_6O_7(OH)_6 \cdot 3(H_2O)$	26
60	01-076-0540	Admontite	$MgO(B_2O_3)_3 \cdot 7H_2O$	73
	01-070-1902	Mcallisterite	$Mg_2((B_6O_7)(OH)_6)_2 \cdot 9H_2O$	38
	01-076-0540	Admontite	$MgO(B_2O_3)_3 \cdot 7H_2O$	48
120	01-070-1902	Mcallisterite	$Mg_2((B_6O_7)(OH)_6)_2 \cdot 9H_2O$	50
	01-073-0638	MgBH*	$MgB_6O_7(OH)_6 \cdot 3(H_2O)$	19
	01-076-0540	Admontite	$MgO(B_2O_3)_3 \cdot 7H_2O$	45
240	01-070-1902	Mcallisterite	$Mg_2((B_6O_7)(OH)_6)_2 \cdot 9H_2O$	35
	01-073-0638	MgBH*	$MgB_6O_7(OH)_6 \cdot 3(H_2O)$	13

\* MgBH: Magnesium Borate Hydrate

From the XRD results obtained it is seen that “01-076-0540” and “01-070-1902” pdf numbered “Admontite” and “Mcallisterite” minerals are formed at all reaction times. 30

minutes after the reaction mcallisterite type of magnesium borate mineral is formed majorly and admontite and MgBH minerals formed also. At the reaction time of 60 minutes all MgBH and partly mcallisterite minerals are changed to admontite minerals. 120 minutes after the reaction some of the admontite minerals are changed to mcallisterite and MgBH minerals oppositely at a reaction time of 60 minutes. After the 240 minutes all minerals scores are decreased, means that their crystalline structures became less crystal than at the reaction time of 120 minutes.

### C. Synthesized Magnesium Borate FT-IR Results

FT-IR spectrums of synthesized minerals are shown in Fig 10.

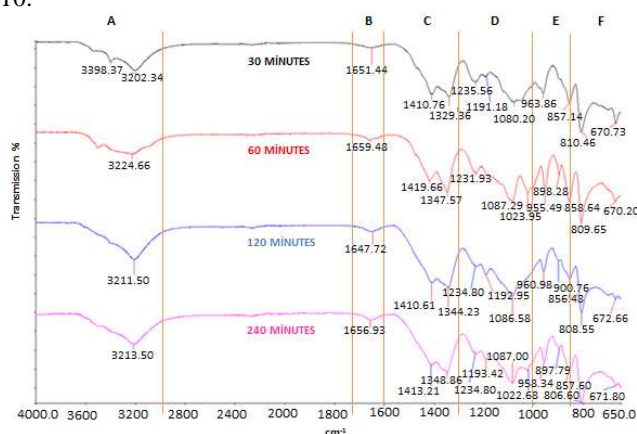


Fig. 10 FT-IR spectrum of synthesized minerals

In the FT-IR spectrums the peak values between  $3398.37$  and  $3202.34$   $cm^{-1}$  that are also seen in the  $B_2O_3$  raw material and these peaks are might be the  $OH^-$  group. At the peaks around  $1650$   $cm^{-1}$  is different peak group from  $B_2O_3$ , can be explained with the water groups in the minerals. The peaks with  $1419.66$  and  $1329.36$   $cm^{-1}$  might be the asymmetric stretching of tri-coordinate boron ( $B_{(3)}-O$ ). The peaks between  $1235.56$  and  $1022.68$   $cm^{-1}$ , might be the asymmetric stretching of tri-coordinate boron ( $B_{(4)}-O$ ). Symmetric stretching of  $B_{(3)}-O$  can be seen between  $693.86$  and  $857.14$   $cm^{-1}$ . Other peaks between  $810.46$  and  $670.20$   $cm^{-1}$  explains the out-of-plane  $OH^-$  bending band and bending of  $B_{(3)}-O$  in the structure.

### D. Synthesized Magnesium Borate RAMAN Results

Raman spectrums of synthesized minerals are shown in Fig. 11. According to Raman spectrums of products, the first peaks in part A also seen on  $B_2O_3$  Raman spectrum that might be the  $OH^-$  groups in the minerals. The peaks of asymmetric stretching of  $B_{(4)}-O$  is seen in part B. In part C, the peak values between  $961.98$  and  $880.68$   $cm^{-1}$  can be explain symmetric stretching of  $B_{(3)}-O$ . The characteristic band of Magnesium boron hydrates is seen at the bands between  $640.17$  and  $635.88$   $cm^{-1}$ , namely  $V_p[B_6O_7(OH)_6]^{2-} / V_p[[B_3O_3(OH)_4]^-]$ , in part D. Peaks of part E, might be explain with the bending of  $B_{(3)}-O$ . Part F includes the peaks of bending of  $B_{(4)}-O$ .



