Determination of the Zinc Oxide and Boric Acid Optimum Molar Ratio on the Ultrasonic Synthesis of Zinc Borates


Abstract—Zinc borates are used as a multi-functional flame retardant additive for its high dehydration temperature. In this study, the method of ultrasonic mixing was used in the synthesis of zinc borates. The reactants of zinc oxide (ZnO) and boric acid (H₃BO₃) were used at the constant reaction parameters of 90°C reaction temperature and 55 min of reaction time. Several molar ratios of ZnO:H₃BO₃ (1:1, 1:2, 1:3, 1:4 and 1:5) were conducted for the determination of the optimum reaction ratio. Prior to synthesis the characterization of the synthesized zinc borates were made by X-Ray Diffraction (XRD) and Fourier Transform Infrared Spectroscopy (FT-IR). From the results Zinc Oxide Borate Hydrate [ZnB₂O₄·3.5H₂O], were synthesized optimum at the molar ratio of 1:3, with a reaction efficiency of 95.2%.

Keywords—Zinc borates, ultrasonic mixing, XRD, FT-IR, reaction efficiency.

I. INTRODUCTION

BORON is not found in nature as an element form, it interacts with carbon and other similar elements and form compounds of boron minerals. Industrial production of high-purity boron is carried out in an expensive and difficult process. Depending on the main elements, they contain boron minerals, calcium borate, magnesium borate, sodium borate, sodium-calcium borate and other borate compounds are divided into five main groups. Zinc borate can be synthesized in a laboratory, although not found naturally. Zinc borate has many application areas ranging from polymers to paints. Different types of zinc borates that are important inorganic hydrated borates can be used as flame and fire retardant and corrosion inhibitor [1], [2]. Depending the contents of zinc and boric oxides, its properties varies and used widely in plastic, rubber, ceramics, paint, wire, electrical insulation, wood applications, cement is commonly used in pharmaceutical industry. In addition, synthetic hydrated zinc borate can be classified as metal borates [3].

II. MATERIALS AND METHODS

A. Raw Materials

Zinc borates were produced by using zinc oxide and boric acid as flame retardant and corrosion. Zinc and boron oxide content to vary depending on their properties and plastic, rubber, ceramics, paint, cable, electrical insulation, wood applications, cement is commonly used in pharmaceutical industry. In addition, synthetic hydrated zinc borate can be classified as metal borates. In ultrasonic irradiation, the chemical effects of ultrasound do not come from a direct interaction of the ultrasonic sound wave with the molecules in the solution. The simplest explanation for this is that sound waves propagating through a liquid at ultrasonic frequencies do so with a wavelength that is significantly longer than that of the bond length between atoms in the molecule. Therefore, the sound wave cannot affect that vibrational energy of the bond, and can therefore not directly increase the internal energy of a molecule.

In this study, using zinc oxide and boric acid a new method of ultrasonic irradiation is used for the synthesis of zinc borates and their effects on the production of the reaction temperature and is investigated. The materials and products are analyzed with X-Ray Diffraction (XRD) and Fourier Transform Infrared Spectroscopy (FT-IR).

B. Ultrasonic Synthesis and Characterizations

Initial experiments were carried out using different ways for producing zinc borate with zinc oxide and boric acid. In this experiment, varied molar ratios of the ZnO(Zn) and H₃BO₃ (H) were tested. Experiment temperature was chosen at 90°C,
and reaction time is set to 40-60 minutes. These parameters were estimated from [6]. For the experiments, between 0.1 and 0.2 mol H$_3$BO$_3$ were dissolved in 20-30 ml pure water in reaction temperature, then different molar ratios of ZnO was added to reactor. After the addition of the raw materials, commercial zinc borate (Zn$_3$B$_6$O$_{12}$·3.5H$_2$O) received from local market in Turkey (in terms of H$_3$BO$_3$, 1% w/w) as added for the better crystallization.

![Fig. 1 Philips PANalytical XRD](image1.jpg)

![Fig. 2 Perkin Elmer Spectrum One FT-IR Spectrometer](image2.jpg)

The raw materials were reacted by ultrasonic irradiation method in a closed temperature-controlled system. At the end of the reaction, each solution was filtered through Whatman blue ribbon filter paper and the crystallized products on the filter paper were washed with the 1-2 L pure water at 60°C to eliminate unreacted boric acid. Then, filtered products dried in the incubator (Ecocell 111, Germany) at 105°C for 1-2 h. Obtained products were characterized by XRD and FT-IR.

### III. RESULTS AND DISCUSSION

#### A. Raw Material Characterization

XRD patterns and results of ZnO, H$_3$BO$_3$ and commercial Zn$_3$B$_6$O$_{12}$·3.5H$_2$O were given in Figs. 3-5 and Table I.

From the XRD analysis of ZnO, it is seen that compound was consist of “01-089-7102” compound that their structural formula is ZnO. H$_3$BO$_3$ and commercial Zn$_3$B$_6$O$_{12}$.3.5H$_2$O were found as, “01-073-2158” coded sassolite (H$_3$BO$_3$) and “00-035-0433” coded zinc oxide borate hydrate, respectively.

![Fig. 3 XRD pattern of ZnO](image3.jpg)

![Fig. 4 XRD pattern of H$_3$BO$_3$](image4.jpg)

![Fig. 5 XRD pattern of commercial Zn$_3$B$_6$O$_{12}$.3.5H$_2$O](image5.jpg)

<table>
<thead>
<tr>
<th>Reference Code</th>
<th>Compound Name</th>
<th>Chemical Formula</th>
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<tbody>
<tr>
<td>01-089-7102</td>
<td>ZnO</td>
<td>ZnO</td>
</tr>
<tr>
<td>01-073-2158</td>
<td>Sassolite</td>
<td>H$_3$BO$_3$</td>
</tr>
<tr>
<td>00-035-0433</td>
<td>Zinc Oxide Borate Hydrate</td>
<td>Zn$_3$B$<em>6$O$</em>{12}$·3.5H$_2$O</td>
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</tbody>
</table>

FT-IR spectrum of ZnO, H$_3$BO$_3$, and commercial Zn$_3$B$_6$O$_{12}$.3.5H$_2$O were given in Figs. 6, 7 and 8, respectively.
According to the FT-IR inorganic library search, H$_3$BO$_3$ was found as: “Boric acid (H$_3$BO$_3$)” with 0.704 score (out of 1) and “AI0031” code.

Commercial Zn$_3$B$_6$O$_{12}$.3.5H$_2$O was not found in the FT-IR inorganic library search, but the boron-oxygen characteristic peaks were observed in the spectrum. The detailed examination will be done at the results section.

### B. Synthesized Products

The XRD results of the synthesized zinc borates were given in Table II.

<table>
<thead>
<tr>
<th>Molar Ratio (Z:N:H)</th>
<th>Reference code</th>
<th>Mineral Name</th>
<th>Mineral Formula</th>
<th>Score</th>
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<tbody>
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<td>1:1</td>
<td>00-035-0433</td>
<td>Zinc Oxide Borate Hydrate</td>
<td>Zn$_3$B$<em>6$O$</em>{12}$.3.5H$_2$O</td>
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<tr>
<td>1:2</td>
<td>00-035-0433</td>
<td>Zinc Oxide Borate Hydrate</td>
<td>Zn$_3$B$<em>6$O$</em>{12}$.3.5H$_2$O</td>
<td>40</td>
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<td>1:3</td>
<td>00-035-0433</td>
<td>Zinc Oxide Borate Hydrate</td>
<td>Zn$_3$B$<em>6$O$</em>{12}$.3.5H$_2$O</td>
<td>65</td>
</tr>
<tr>
<td>1:4</td>
<td>00-035-0433</td>
<td>Zinc Oxide Borate Hydrate</td>
<td>Zn$_3$B$<em>6$O$</em>{12}$.3.5H$_2$O</td>
<td>63</td>
</tr>
<tr>
<td>1:5</td>
<td>00-035-0433</td>
<td>Zinc Oxide Borate Hydrate</td>
<td>Zn$_3$B$<em>6$O$</em>{12}$.3.5H$_2$O</td>
<td>58</td>
</tr>
</tbody>
</table>

XRD scores of the zinc borates synthesized from zinc oxide and boric acid in same temperature and same reaction time (90°C and 55 minutes) is shown in Table II. Synthesized zinc borate from zinc oxide and boric acid was found as ‘zinc oxide borate hydrate (Zn$_3$B$_6$O$_{12}$.3.5H$_2$O)’ with powder diffraction file number (pdf number) of ‘00-035-0433’. From the XRD scores of synthesized zinc borates the best molar ratio is seen as 1/3. Obtained zinc borates coded as ‘set code – mole ratio - reaction temperature – reaction time’ which the sets codes are ‘Z-H 1:1 90-55’. XRD patterns of the selected zinc borates were given in Fig. 9.

FTIR spectra of the products are given in Fig. 10. Bending vibrations of the H-O-H band are a little bit observed between 1500 and 1420 cm$^{-1}$. The existence of the band between 1420-1339 cm$^{-1}$ is showed to the asymmetric stretching vibrations of trihedral (BO$_3$) borate groups. The peaks in the range between 1175-1058 cm$^{-1}$ is observed to the asymmetric stretching vibrations of tetrahedral (BO$_4$) borate groups. Bending of symmetric stretching vibrations of trihedral (BO$_3$) borate groups are showed between 1058 and 924 cm$^{-1}$.The peaks in the range between 876-859 cm$^{-1}$ is showed to the symmetric stretching vibrations of tetrahedral borate groups. The peak observed between 859-659 cm$^{-1}$ indicates the plane bending vibrations of trihedral borate groups.
In this study, zinc borates were synthesized optimum at the molar ratio of 1:3, with a maximum reaction efficiency of 95.2%. At the future studies, reaction time and the reaction temperature changes will be investigated in the synthesis of zinc borates and may be produced with different zinc and boron materials.

### REFERENCES


